

INTRODUCTION TO  
CHAPTER SEVENTEEN

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As the massive nuclear process which was used to create this Earth came to an end, the Earth was left with a sterile environment. This environment was deadly to life, because of all the dangerous remnants of the nuclear and chemical processes which were in ferocious action during the early stage of Creation. This hazardous situation existed during Day One, and possibly even Day Two of the Creation week. But God had an important plan for this Earth. He created it to be inhabited by many special forms of life. It was meant to be a safe and productive environment for humans, animals, and vegetation, among other things. Therefore, the face of the Earth needed to be completely decontaminated.

This chapter investigates a facet of the immense decontamination process which it appears God set in motion to purge the face of the Earth of the many detrimental substances. This process appears to have occurred prior to the Creation of the first landmass on Day Three, and the many forms of vegetation which it contained. This chapter examines a number of physical and chemical processes which were used to form a number of Earth-compounds. These processes are put into relatively simple, easy to understand language. By these relatively simple, yet very effective processes, God rapidly cleaned up the Earth, as He prepared to make it habitable.

During the massive clean-up process, a number of new compound-layers were formed upon the 'basement rocks' of the Earth's crust. These layers are commonly found below the surface of the Earth, except where substantial geologic upheaval has thrust these layers upward, into the view of all mankind. To readily grasp the concepts in this chapter, the Reader will need to take the processes used in the simple examples, and mentally intensify these processes into an immense, world-class scale. They must be intensified to that scale which God commonly uses. Only then can anything near justice be done to the mighty acts which God performed at Creation.

May due glory, honor, and praise be given to God the Father, and to His beloved Son, our Lord and Saviour, Jesus Christ. May God be rightly worshiped for all of the mighty works which He has done. May He be worshiped by our willing obedience to Him.

# Chapter 17: THE CLEANUP

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## EVEN SO COME

The prophet cries to the Lord: *“Oh that thou wouldest rend the heavens, that thou wouldest come down, that the mountains might flow down at thy presence. As when the melting fire burneth, the fire causeth the waters to boil, to make thy name known to thine adversaries, that the nations may tremble at thy presence!”*<sup>1</sup> Amen, and amen! *“Even so come, Lord Jesus!”*<sup>2</sup>

## PRESENT STUDY

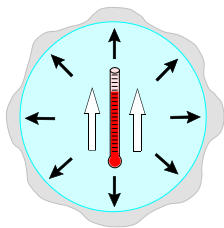


This chapter examines only those processes which were occurring within the worldwide ocean: that ocean which was forming on the first day of Creation. This chapter examines one of the most likely, logical, and practical processes by which God cleaned up the worldwide acidic ocean of harmful chemicals.

This particular study only examines a few processes which are readily obvious to the writer. He is certain that there were numerous other processes in action at the same time, to one degree or another.

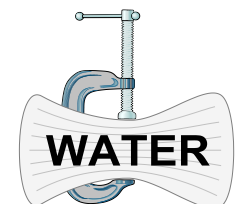
First, let us examine what happens to superheated water which is under extreme pressure. This would have been the state of the Earth’s first waters.<sup>3</sup> Afterward, we will move on to other matters.

## ULTIMATE RELEASE



The critical point of a substance is that temperature or pressure at which a well-defined change takes place in its physical state. The critical temperature of a liquid is that particular temperature at which it will convert into a gas no matter how much external pressure is exerted on it.<sup>4</sup>

The critical temperature for water is reached at 705.6°F (374.2°C). At this temperature, a total of 218 atmospheres of pressure are required to cause water to remain in its liquid state.<sup>5</sup> A pressure of 218 atmospheres is equal to about 3,204.6 pound per square inch (2,253,154 kgs/sq m).



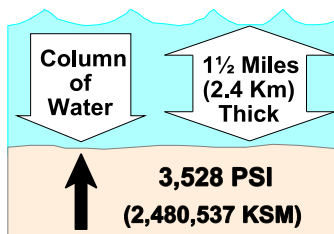
**218 Atmospheres  
Pressure**

Above the critical temperature of 705.6°F (374.2°C), water cannot exist in its liquid state. Above this temperature, it will instantly flash into its gaseous state. This change of state will occur no matter how much additional pressure is put to bear upon the water.

## DEEP WATERS

If the current total volume of the ocean waters were averaged out over the face of the whole Earth, the solid portion of the Earth would be submerged under a layer of water about 1½ miles thick (about 7,920 feet or 2,414 m).<sup>6</sup>

## GREAT WATER PRESSURE



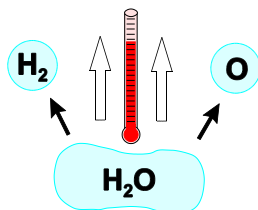
In the ocean, the water pressure increases by about one atmosphere (approximately 14.7 pounds per square inch) for each additional 33 feet (10 m) of depth attained.<sup>7</sup> The pressure at the bottom of a 7,920 foot (2,414 m) thick layer of water would be equal to approximately 3,528 pounds per square inch (2,480,537 kgs/sq m).

## SIMPLE FACT

On the first day of Creation, there was a very deep ocean of boiling water which covered the face of the whole Earth. The above-noted-pressure at the very bottom of a worldwide ocean is only about 323 pounds per square inch (227,101 kgs/sq m) greater than the pressure required to keep water liquid at only its critical temperature. By all the evidence, it would be impossible for this pressure to keep the water in a liquid state above the critical temperature of 705.6°F (374.2°C).

From the evidence which is soon examined, it appears that any water in near proximity to the molten surface of the Earth would have flash boiled, and converted into steam. The layers of water above would have less weight bearing down upon them, therefore, they would also be boiling vigorously.

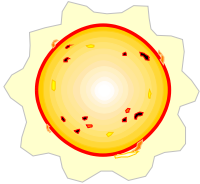
## THERMAL DECOMPOSITION



At a temperature of 4,892°F (2,700°C), about 11 percent of the water molecules in a vapor are decomposed into the component gases of hydrogen and oxygen.<sup>8</sup> As the temperature is increased, so also is the percentage of water molecules which are decomposed. At a temperature of 6,332°F (3,500°C), about 30 percent of the water molecules in a sample are decomposed into hydrogen and oxygen.<sup>9</sup>

## SIMPLE COMPARISON

Friend, it is believed that God the Father and His Son created the inner core of the Earth by an extremely large and intense nuclear reaction. In our solar system, the only thing which we have for ready reference to examine the temperatures attained during a large-scale nuclear reaction is the sun.



The scientists claim that the core temperature of the sun is about 29,000,000°F (16,111,093°C). The surface of the sun is believed to be at a radial distance of about 432,500 miles (695,893 km) from the nucleus of the sun. At this great distance from the nucleus, the surface of the sun is claimed to have a temperature of 9,300°F (5,149°C), or more.<sup>10 and 11</sup>

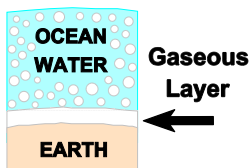
## EARTH INFO

The scientists believe that the present temperature at the inner-core of the habitable Earth is as high as 12,630°F (7,000°C).<sup>12 and 13</sup> The surface of the Earth is currently located at a radial distance of only about 3,963 miles (6,376 km) from the superheated, plasmatic nucleus of the Earth.<sup>14</sup>

During the initial stage of nuclear Creation, the outer temperature of the superheated Earth appears to have been extremely high. It would have been far in excess of that temperature which is required to boil water. It would have been far above the critical temperature of water. It may have even been high enough at various locations (or possibly even worldwide) to thermally decompose any water molecules which were in the vicinity of the outer surface of the Earth.

## A BOUNDARY LAYER

In the beginning, an extremely thick, superheated layer of gases may have been directly in contact with the forming surface of the Earth. Within this gaseous zone may have been great quantities of free hydrogen and oxygen, which were igniting and forming a layer of superheated steam.



On Day One of Creation, the condensing ocean of water may have initially been separated from the surface of the molten Earth by a great distance. There would have been nothing but a very thick gaseous layer between the Earth and the vigorously boiling and condensing worldwide ocean above.

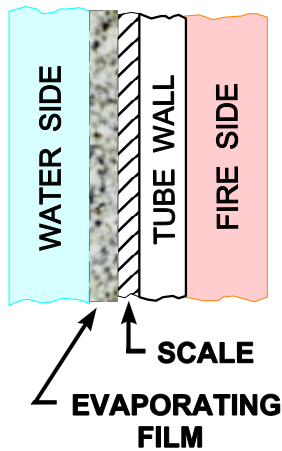
## AN EXAMPLE

For a general example of the above phenomenon, let us examine a modern feature in which a similar process occurs, but only on a much smaller scale. Let us look at a large water-boiler. We will consider the type which is commonly used for producing electrical power in large central stations.

## GASEOUS FILM

In a boiler, heat from a fuel source is applied to the boiler vessel. In a central station, the boiler vessel usually consists of metal tubes which are filled with water. The applied heat passes through the metal of the boiler vessel, and into the water contained therein. In a well-designed boiler system, there will be an evaporating film-layer between the metal of the boiler vessel and the water.<sup>15</sup> Within the region of this evaporating film-layer, certain types of buildups may occur.

## DEPOSITS OF SCALE



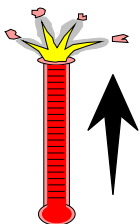
Deposits of scale and sludge may form on the heated surfaces of boiler tubes where they contact the water during the process of converting water into steam. Any chemical substances which are dissolved in the water will become concentrated and form a coating on the evaporating surface.<sup>16</sup> Further heating may then convert this layer of foreign substance into a hard scale.

As the temperature is increased, the types of substances which separate from the water to form scale tend to become less soluble in water. In this situation of increased temperature, the deposition rate of the scale-forming substances on the evaporative surface will also greatly increase.<sup>17</sup> In other words, the higher the temperature, the faster the scale will increase in thickness upon the evaporative surface.

## BAKED SCALE

On the first day of Creation, a superheated gaseous layer existed between the scorching-hot Earth and the boiling waters of the condensing ocean. The minerals from the Earth did not readily pass to the other side of this gaseous barrier. On the surface of the Earth, a thick salt layer was forming, and rapidly baking into a hard rock. This process resulted in thick halite deposits upon this Earth.<sup>18</sup>

## RAGING WATERS

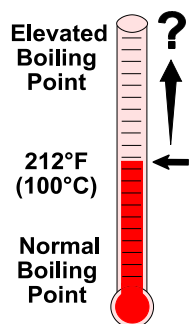


Friend, during the immense nuclear reaction at the Creation of this Earth, an exceeding great amount of heat was generated. A portion of this heat was contained within the primitive waters which began to condense and surround the basic 'blob' of the Earth. As the heat from below entered into the condensing waters, it kept them boiling vigorously and continually vaporizing. This rapidly accomplished the halite process shown above!

Because of the processes by which this Earth was created, the ejected residues of the nuclear reaction were mixed up in the forming, worldwide waters. These waters may have appeared as an immense, soupy, boiling mud puddle. Let us now examine the situation within these waters.

## ABOVE NORMAL

Pure water boils at 212°F (100°C) when at a sea-level atmospheric pressure of about 14.7 pounds per square inch (10,336 kgs/sq m). When a non-volatile substance is added to this same water, its boiling point changes.



The addition of a non-volatile solute to water causes its boiling point to increase. As more solute is added, the boiling point of the water is raised even further. In other words, when there is a great concentration of solute in water, its temperature must be increased greatly to bring it to a boil.<sup>19</sup>

Friend, during the first few days of Creation, because of the prevailing conditions, it appears that there may have been an extremely high concentration of non-volatile materials in the primitive ocean. Exactly how high, then, could have been the boiling point of these ancient waters?

## TEMPERATURE FACTOR

On the first few days of Creation (as the Earth was cooling), the high heat and material flow into the primitive waters allowed a large excess of salts to be in solution. It is a known fact that for a given amount of solvent (even one such as ordinary water) the amount of a given solid substance which will dissolve into it normally increases as the temperature is increased.<sup>20 and 21</sup>

In other words, as the temperature goes up, the solubility of a substance in a solvent usually goes up. There are relatively few salts whose solubility does the opposite, and actually decreases as the temperature is increased.<sup>22 and 23</sup>

## INCREASED SOLUBILITY

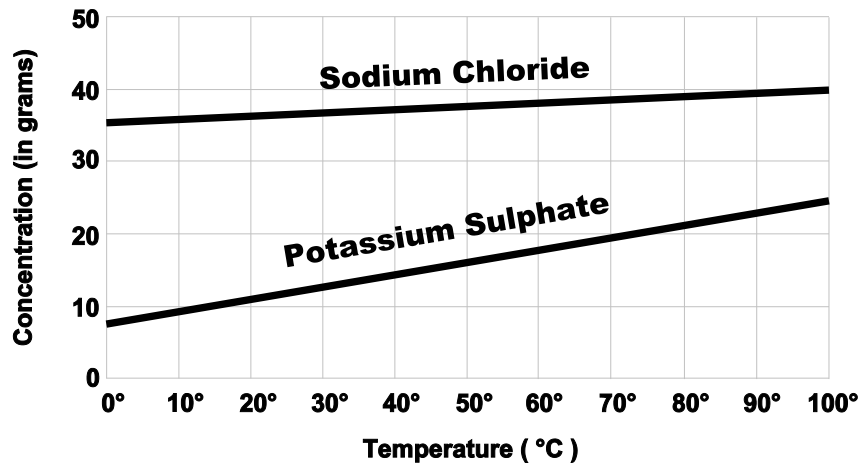
Sodium chloride is the most plentiful salt found in seawater. It makes up about 78 percent of the total salt content within the water. About 35 grams of sodium chloride will dissolve in a 100 gram sample of water at 32°F (0°C).<sup>24 and 25</sup>

When the temperature of the 100 gram water sample is raised to 212°F (100°C), about 40 grams of sodium chloride will dissolve in the same water sample.<sup>26 and 27</sup> This is about a 14 percent increase in solubility when the temperature of the water is raised from its freezing point to its boiling point.

## GRAPHIC EXAMPLE

The left hand column in the graph below indicates the concentration (in grams, within a 100 gram sample of water) for just two of those salts which are contained within the oceans. The salts shown are sodium chloride and potassium sulphate. The bottom row indicates the temperature of the water sample, in degrees Centigrade.

### A MARKED DIFFERENCE



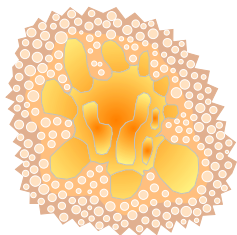
The lower heavy line (on the graph above) indicates the solubility increase for potassium sulphate, compared to that of sodium chloride (common salt). The potassium sulphate makes up about 2.5 percent of the total salt content in the current, average ocean waters.

Shown graphically above, in a 100 gram sample of water, roughly 8 grams of potassium sulphate will readily dissolve at 32°F (0°C). When the temperature of this same water sample is increased to 212°F (100°C), which is the normal boiling point of water at sea-level, about 24 grams of potassium sulphate will readily dissolve.<sup>28</sup> As pressures greatly increase (such as at the bottom of an ocean), the boiling point also greatly increases, as does the solubility of dissolved substances.



The above information shows a three-fold increase in solubility for potassium sulphate when the temperature of the water sample is raised from its freezing point to just its sea-level boiling point. A much greater volume of potassium sulphate may have been dissolved within the highly pressurized, superheated waters at Creation.<sup>29</sup> This may have worked to create an even greater precipitation of salts as the waters began to cool.

## IN THE BEGINNING



Once again, when God first created the basic 'blob' of this Earth, there was a deep, vigorously boiling ocean of water which covered the whole face of the Earth. The composition of the waters of this ocean was far different from that which is found in the oceans of today.

The waters of the original worldwide ocean were extremely acidic. They were full of hydrochloric acid, and many other acids and toxic substances. The acidic nature of these waters was a result of all the nuclear and chemical processes involved with the Creation of the basic 'blob' of the Earth.

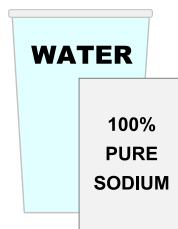
The vast ocean of acrid waters desperately needed to be rendered safe for sea life. In just four days, it was God's plan to create fishes, and many other types of sea creatures. It was intended that they should live in the waters which would span across the face of the Earth. Full well knowing the necessary order for His great works, God rapidly began preparing the waters for living things.

## NEUTRALIZATION



As God the Father and His beloved Son began to prepare the waters of the vast ocean for the myriad of living creatures which would soon be created, an important substance needed to be prepared. We of today would call this substance a chemical base. This chemical base, when thoroughly mixed into the acrid waters of the Sea (in the proper proportions) would make the waters safe. Because the ocean waters were vigorously boiling, the chemical base readily and thoroughly intermixed with all of the acidic water.

## BASE EXAMINATION



What was this chemical base which God used to clean the acids out of the vast, worldwide ocean? For the most likely answer, please examine the Periodic Table and Specific Gravity charts in the preceding chapter. With a little knowledge in chemistry, it appears that a very likely candidate for producing a good chemical base is the lightweight metal called sodium.

Sodium (chemical symbol: Na) has a specific gravity of 0.97. It is a highly reactive metal. Because of its low specific gravity, the sodium rose upward through the layers of molten materials. Turbulence caused mixing of the elements, and the creation of numerous compounds. It also caused the sodium to come into contact with water vapors. As the sodium reacted with these water vapors, the neutralizing chemical-base was formed.



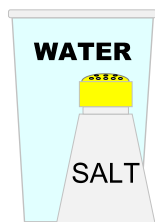
## A COMMON BASE



A solution of sodium hydroxide (NaOH) is a strong chemical base. It is commonly called caustic soda, or just simply lye.<sup>30</sup> A sodium hydroxide solution was an excellent chemical base to use during the Creation cleanup process. Indeed, God is extremely wise in all of those things which He does! He was, as always, extremely wise in the use of sodium hydroxide! This particular chemical base is very stable. It does not decompose, even when heated to a red-hot heat.<sup>31</sup> This characteristic would allow this base to readily operate within the vigorously boiling waters of the early Sea.

When sodium hydroxide is properly mixed with hydrochloric acid (HCl), the acid is completely neutralized.<sup>32</sup> During those first few days of Creation, the sodium hydroxide was rapidly intermixed with those vigorously boiling and churning waters of the vast, acidic, primitive ocean. Because of the perfect intermixing within these turbulent waters, the neutralization process occurred very rapidly. This process was also very complete.

## SALT



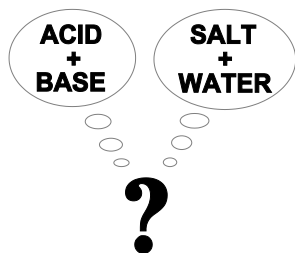
In science, the term “salt” is commonly applied to those substances which are created by the reaction of an acid with a base (although this is not the only manner whereby salts are created).<sup>33</sup> When sodium hydroxide was mixed with the acidic ocean waters, the end products of this neutralization process were common table salt (NaCl) and ordinary water (H<sub>2</sub>O).<sup>34</sup> As the neutralization process was nearing its completion, vast quantities of salt began to come out of solution and precipitate to the bottom of the Sea.

## MORE METHODS

There are a number of other methods by which a salt is produced. In one method, a metal (such as sodium, etc.) may react directly with an acid to produce a salt. In another method, two elements (such as sodium and chlorine) are also able to combine under the proper circumstances, to produce common table salt.<sup>35</sup>

Friend, there truly were numerous nuclear and chemical processes in operation at the Creation of this Earth. The writer believes that a number of these processes worked together to rapidly produce the extremely thick layers of rock salt which are found buried beneath the continents today. Please note: The sole purpose of this brief study is simply to point you in a particular direction, one which you may never have considered before (and one which the scientists apparently have neglected to tell you). The way is now prepared for you to investigate this matter further, as God guides you.

## CHEMICAL PUZZLE

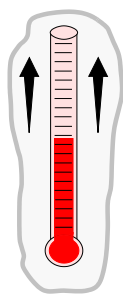


When the neutralization process of the worldwide Sea was complete (possibly by the third day of Creation, and for sure by the fifth day of Creation, before all the fishes and other creatures were placed into the Sea), the products which remained were mainly water and sodium chloride salt. It should be noted that a chemist can examine these end products, and will be unable to tell if this system of water and salt was created by simply dissolving a salt in the water, or by mixing an acid with a chemical base.<sup>36</sup>

Friend, for the above reason, the scientists may have completely overlooked the truth of that which occurred at Creation. They truly have missed the simple procedure and the exceedingly great magnitude with which God the Father and his beloved Son, Jesus Christ, worked as all things were created. Truly, they have missed seeing the magnificent beauty of God. Friend, God has hidden these special things from the wise and prudent of this world, and has given them unto you!

Friend, the final products of neutralization are indeed very indistinguishable from the final products of other chemical methods for producing salts. Please, let us not too quickly, or possibly unnecessarily fault any misguided scientists for this potential oversight. May this oversight cause all persons to clearly understand their need for the complete Word and Spirit of God! May we all, in the end, come to an even greater understanding of the Lord God Almighty, and His beloved Son, our Lord and Savior Jesus Christ, by whom we were magnificently created. May we come to understand the great beauty and awesome power with which God works.

## HEAT OF REACTION



As shown in the preceding sections, when a strong acid is properly mixed with a strong chemical base, a neutralization process occurs. During this neutralization reaction, a rather large amount of heat is rapidly liberated from the solution.

For those with an understanding of science, in a reaction between hydrochloric acid and sodium hydroxide, the quantity of heat liberated amounts to almost 13.6 kcal per mole of substance.<sup>37</sup>



Another source states the particular quantity of heat produced (during this neutralization process between hydrochloric acid and sodium hydroxide) in a somewhat different manner. They also give a somewhat different overall heat-value. They state that in the above exothermic neutralization reaction, for every gram-mole of water produced, -13,800 calories of heat are released.<sup>38</sup>

## TERM DEFINITION

In common scientific terms, a calorie is that specific quantity of heat which will cause an increase in the temperature of one gram of water by 1°C.<sup>39</sup> The claimed -13,800 calorie release of heat per each gram-mole of water (produced in the above neutralization process) amounts to a considerable heat output from the former worldwide ocean which existed during the time of Creation. The great amount of heat produced during the neutralization process may have caused the primitive ocean to suddenly flash-boil. Water and other superheated materials would have been thrown into the primitive atmosphere in every direction at this sudden release of energy!<sup>40</sup>

## FIRST CLEAN

In the preceding pages, the use of a sodium hydroxide solution at Creation was shown. It is of interest to note that a layer of sodium hydroxide solution readily absorbs carbon dioxide out of the adjoining atmosphere.<sup>41</sup> By the initial absorption process, a great volume of carbon dioxide was removed from the primitive atmosphere. This greatly reduced the cleanup which the plant-life would have to do to the air, once they were created on day three of Creation.

During the carbon-dioxide absorption-process, a chemical conversion takes place, the end product being sodium carbonate ( $\text{Na}_2\text{CO}_3$ ). This sodium carbonate is insoluble in the sodium hydroxide solution, therefore, it precipitates out of solution and forms a bottom-layer.<sup>42</sup> On the first days of Creation, a thick layer of sodium carbonate was formed at the bottom of the worldwide Sea. These layers eventually became concentrated and compressed, thereby converting into deposits of trona. The largest known trona deposit in this world is found in the state of Wyoming, USA.<sup>43</sup>

## MORE LAYERS

A solution of sodium carbonate may become mixed with a solution of a beryllium salt (a good example of a beryllium salt is  $\text{BeSO}_4$ ). When this occurs, a basic beryllium carbonate with the formula  $(\text{Be}_2\text{O})\text{CO}_3$  is precipitated.<sup>44</sup> This precipitate may form into economically valuable deposits from which beryllium may be recovered.

When a solution of sodium carbonate becomes mixed with a solution of a magnesium salt (a good example being  $\text{MgSO}_4$ ), a basic magnesium carbonate with the formula  $\text{Mg}_2(\text{OH})_2\text{CO}_3 \cdot 3\text{H}_2\text{O}$  is precipitated.<sup>45</sup> Another source gives the previous precipitate the simplified chemical formula of  $\text{MgCO}_3$ .<sup>46</sup> These precipitates may form valuable deposits from which may be obtained metallic magnesium.

We now end this basic study of a few of those interactions which appear to have been used to create valuable mineral deposits, as found worldwide. This study is simply a very small “stepping stone.” It may point the Reader in the proper direction for further, independent research.

## TURBULENT SEAS

Once again, on the first day of Creation, the worldwide ocean was in a great, rolling boil. Boiling is actually a result of localized evaporation. It is the formation of a gaseous bubble within the heated substance. Evaporation (including boiling) is a cooling process.<sup>47</sup> In the process of boiling, the bubbles transfer excess heat from the evaporative surface to the zone above the heated liquid substance. On the first day of Creation, there was no protective atmosphere around the primitive 'Earth-blob.' In this situation, the heat energy was rapidly drawn off the Earth into the deep cold of space. This, in turn, helped to cause an extreme, rapid cooling (almost a quick-freeze) of the Earth.

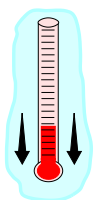
## RAPID COOLING



Friend, now it is time to put all of the foregoing information toward a practical use! Let us use it as we consider the following questions. What effect would a rapid reduction in water temperature have had on the immense volume of salts and other minerals dissolved in the primitive ocean, during those first few days of Creation? Would a rapid drop in temperature have caused an enormous precipitation of many different types of dissolved salts, plus other sedimentary-types of materials?

## STANDARD PROCESS

When an agitated solution begins to cool, the excess substances in solution begin to crystalize. As these substances crystalize, they precipitate out of solution. This precipitation continues until a new solubility equilibrium is established at the new temperature.<sup>48</sup> If the temperature continues reducing, the dissolved salts may also continue precipitating out of solution.



During the early days of Creation, the heat within the waters of the original ocean was swiftly drawn off into the deep cold of space. Once again, there was no atmosphere yet to prevent this heat escape from occurring. Therefore, an extremely rapid drop in water temperature occurred. This drastic drop in water temperature reduced the solubility of the salts contained therein, and ultimately caused an extremely large precipitation of salts.

## A GREAT PRECIPITATION

As the immense volumes of salt precipitated out of solution, the layers of individual and intermixed salts (and other minerals) at the bottom of the worldwide ocean grew thicker and thicker. The vast salt deposits initially formed a layer at the heated floor of the primitive, worldwide ocean.

## ROCK SALT

Pure water has a specific gravity of 1.00. The average sample of seawater has a specific gravity of 1.025.<sup>49</sup> The common salt, halite, has a specific gravity of 2.17.<sup>50</sup> This means that halite is more than twice as dense as seawater. In simple layman's terms, a given volume of halite will weight more than twice as much as a equal volume of seawater.

## THICKENING DEPOSITS

As the salt layer grew thicker at the heated bottom of the worldwide ocean (at Creation), the salt in the lower portion of this layer began to compress from the weight of the salt above it. In the end, the lighter-weight water was boiled and squeezed out.<sup>51</sup>

During the process of deposition and compression (by various means), thick layers of rock salt began to form above those which were already created by the simple, but powerful chemical reaction between the chlorine gas and the elemental sodium on Day One of Creation. These layers which were formed later, were likewise baked solid by the heat which radiated from the Earth below.

By the process just shown, and by a number of other chemical processes, God prepared the primitive ocean waters for the numerous living things which He would place therein a few days later. By the above processes, many mineral layers were also formed at the bottom of the ancient Sea.<sup>52</sup>

## CURRENT COMPOSITION

By weight, the average ocean waters currently contain about 3.44 parts of mineral matter in solution for every 100 parts of seawater. The included mineral matter is in the form of salts. A list of the salts, and the percentages of each type in relation to the total salt content, is as follows:<sup>53</sup>

Sodium Chloride	-----	77.758
Magnesium Chloride	-----	10.878
Magnesium Sulphate	-----	4.737
Calcium Sulphate	-----	3.600
Potassium Sulphate	-----	2.465
Magnesium Bromide	-----	0.217
Calcium Carbonate	-----	0.345
Total Percentage Of Salt Content	-----	100.000

Besides the above ingredients in the water of the oceans, there is also a large quantity of included nitrogen, oxygen, and carbonic acid gases. Many of the metallic elements listed on the Periodic Table are also found dissolved in seawater.<sup>54</sup>

## ANCIENT TESTIMONY

Friend, let us now look unto the Lord God Almighty who, by His beloved Son, Jesus Christ, magnificently created the Earth and Sea, and all that is therein. Let us look further into the Word of God to find written words of wisdom.

The psalmist cries: *“The LORD reigneth, he is clothed with majesty; the LORD is clothed with strength, wherewith he hath girded himself: the world also is stablished, that it cannot be moved.*

*“Thy throne is established of old: thou art from everlasting. The floods have lifted up, O LORD, the floods have lifted up their voice; the floods lift up their waves.*

*“The LORD on high is mightier than the noise of many waters, yea, than the mighty waves of the sea. Thy testimonies are very sure: holiness becometh thine house, O LORD, for ever.”<sup>55</sup>, see also <sup>56</sup>*

To God be the glory, both now and forever. Amen.

Friend, so ends another Chapter in this FIRST EDITION of  
*Listen To The Earth, Volume One, THE CREATION*, by *David E. Sakrison and Friends*  
in 34 Chapters, plus README, Preview, Start, and End files  
with References following each Chapter

## REFERENCES

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1. *The Holy Bible*, Book of Isaiah, chapter 64, verses 1-2.
2. *The Holy Bible*, Book of Revelation, chapter 22, verse 20.
3. This would have been the state especially at the very bottom of the very deep waters which covered the face of the primitive Earth.
4. *Microsoft Encarta 98 Encyclopedia*, "Critical Point."
5. *Microsoft Encarta 98 Encyclopedia*, "Critical Point."
6. *The Flood*, Alfred M. Rehwinkel, 1951, page 124.
7. *The 1998 Grolier Multimedia Encyclopedia*, "Dead-sea Life."
8. *Concepts of General Chemistry*, McLellan / Day / Clark, 1966, page 438.
9. *Chemistry*, James V. Quagliano, 2<sup>nd</sup> Edition, 1963, page 236.
10. *Microsoft Encarta 98 Encyclopedia*, "Sun."
11. *Astronomy*, June 1998, "The Blinking Sun Gets Hot," Anonymous, pages 26-28.
12. *Microsoft Encarta 98 Encyclopedia*, "Earth."
13. *Mantle Plumes and Their Record in Earth History*, Kent C. Condie, 2001, pages 4-5.
14. See: *The 1998 Grolier Multimedia Encyclopedia*, "Earth, size and shape of."
15. See: *Combustion Engineering*, Glenn R. Fryling, Revised Edition, 1967, page 6-4, see Figure 6-2.
16. *Combustion Engineering*, Glenn R. Fryling, Revised Edition, 1967, page 8-10.
17. *Combustion Engineering*, Glenn R. Fryling, Revised Edition, 1967, page 8-10.
18. The deposits of many other minerals also took place at this same time. These deposits formed layers within the strata of the Earth.
19. *Chemistry*, James V. Quagliano, 2<sup>nd</sup> Edition, 1963, page 379.
20. *Smith's Introductory College Chemistry*, James Kendall, 1931 Edition, page 140.
21. *Chemistry*, James V. Quagliano, 2<sup>nd</sup> Edition, 1963, page 365.
22. *Smith's Introductory College Chemistry*, James Kendall, 1931 Edition, page 140.
23. *Chemistry*, James V. Quagliano, 2<sup>nd</sup> Edition, 1963, page 365.
24. See: *Smith's Introductory College Chemistry*, James Kendall, 1931 Edition, page 141.
25. See: *Chemistry*, James V. Quagliano, 2<sup>nd</sup> Edition, 1963, page 367.



26. See: *Smith's Introductory College Chemistry*, James Kendall, 1931 Edition, page 141.
27. See: *Chemistry*, James V. Quagliano, 2<sup>nd</sup> Edition, 1963, page 367.
28. See: *Smith's Introductory College Chemistry*, James Kendall, 1931 Edition, page 141.
29. It should be noted that the pressure at the bottom of the worldwide ocean may have been about 3,528 pounds per square inch (2,480,537 kgs/sq m), rather than the standard sea-level pressure of 14.7 pounds per square inch (10,336 kgs/sq m) found at the ocean's surface. This great pressure at the bottom of an ocean (which was being superheated from below) would have increased its boiling point to a great value. When the effect of the non-volatile solids is factored in, the boiling point of the water at the bottom of the worldwide ocean could have been extremely high.
30. *The 1998 Grolier Multimedia Encyclopedia*, "Salt (chemistry)."
31. *Chemistry*, James V. Quagliano, 2<sup>nd</sup> Edition, 1963, page 679.
32. *The 1998 Grolier Multimedia Encyclopedia*, "Salt (chemistry)."
33. *Microsoft Encarta 98 Encyclopedia*, "Salt (compound)."
34. *The 1998 Grolier Multimedia Encyclopedia*, "Salt (chemistry)."
35. *The 1998 Grolier Multimedia Encyclopedia*, "Salt (chemistry)."
36. *The 1998 Grolier Multimedia Encyclopedia*, "Neutralization."
37. *Concepts of General Chemistry*, McLellan/Day/Clark, 1966, pages 200-201.
38. *Chemistry*, James V. Quagliano, 2<sup>nd</sup> Edition, 1963, pages 462-463.
39. *Chemical Technicians' Ready Reference Handbook*, Shugar/Shugar/Bauman, 1973, page 370.
40. This sudden release of energy may have also triggered an equally sudden, immense cascade of materials precipitating out of the solution.
41. *Chemistry*, James V. Quagliano, 2<sup>nd</sup> Edition, 1963, page 679.
42. See: *Chemistry*, James V. Quagliano, 2<sup>nd</sup> Edition, 1963, page 679.
43. *The 1998 Grolier Multimedia Encyclopedia*, "Wyoming."
44. *Chemistry*, James V. Quagliano, 2<sup>nd</sup> Edition, 1963, page 695.
45. *Chemistry*, James V. Quagliano, 2<sup>nd</sup> Edition, 1963, page 695.
46. *Microsoft Encarta 98 Encyclopedia*, "Magnesium."
47. *Conceptual Physics*, Paul G. Hewitt, 6<sup>th</sup> Edition, 1989, page 294.
48. *Concepts of General Chemistry*, McLellan / Day / Clark, 1966, page 250.
49. *The 1998 Grolier Multimedia Encyclopedia*, "Density."
50. *Microsoft Encarta 98 Encyclopedia*, "Salt (compound)."

51. The layers of other sedimentary materials which formed on top of the salt deposits would have also put added pressure to bear upon the salt. This may have helped accelerate the removal of water out of these salt deposits, and their conversion into rock salt.
52. These mineral layers were buried under a layer of relatively loose sediments.
53. From: *Geology*, Chamberlin and Salisbury, 1909, volume 1, page 324.
54. *Geology*, Chamberlin and Salisbury, 1909, volume 1, page 325.
55. *The Holy Bible*, Book of Psalms, chapter 93, verses 1-5.
56. It appears that the basic ball of the Earth may continue, but the outer surface will be destroyed in an intense fire, generally purifying it, that it may be renewed. In the *Holy Bible*, the Book of Psalms, chapter 102, verses 25-27, the psalmist proclaims to God: "*Of old hast thou laid the foundation of the earth: and the heavens are the work of thy hands. They shall perish, but thou shalt endure: yea, all of them shall wax old like a garment; as a vesture shalt thou change them, and they shall be changed: but thou art the same, and thy years shall have no end.*" Friend, a vesture is the outer covering. The vesture of the Earth is the place where the works of mankind reside.  
Considering this matter further, in the *Holy Bible*, the Book of II Peter, chapter 3, verses 10-13, the apostle Peter declares: "*But the day of the Lord will come as a thief in the night; in the which the heavens shall pass away with a great noise, and the elements shall melt with fervent heat, the earth also and the works that are therein shall be burned up. Seeing then that all these things shall be dissolved, what manner of persons ought ye to be in all holy conversation and godliness, looking for and hasting unto the coming of the day of God, wherein the heavens being on fire shall be dissolved, and the elements shall melt with fervent heat? Nevertheless we, according to his promise, look for new heavens and a new earth, wherein dwelleth righteousness.*"  
Friend, the elements shall melt. The Earth and the works that are upon it shall be burned up. All things shall be dissolved. But what form will they then take? Will the new Earth (wherein dwelleth righteousness) be based upon a completely new outer region, which replaces this polluted Earth, or will the new Earth actually be another orb, located at some other completely different position in space? Friend, there is much "food for thought" here, and a need for much prayerful and diligent searching of the Holy Scriptures. Only then will we be able to receive the whole truth for ourselves.